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Chemistry Class - 11: Chapter – 5. States of Matter Part – 1

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I. Multiple Choice Questions (Type-I)

Question: 1

A person living in Shimla observed that cooking food without using pressure cooker takes more time. The reason for this observation is that at high altitude:

- (i) Pressure increases
- (ii) Temperature decreases
- (iii) Pressure decreases
- (iv) Temperature increases

Answer: (iii)

Question: 2

Which of the following property of water can be used to explain the spherical shape of rain droplets?

- (i) Viscosity
- (ii) Surface tension
- (iii) Critical phenomena
- (iv) Pressure

Answer: (ii)

Question: 3

A plot of volume (V) versus temperature (T) for a gas at constant pressure is a straight line passing through the origin. The plots at different values of pressure are shown in Fig. 5.1. Which of the following order of pressure is correct for this gas?

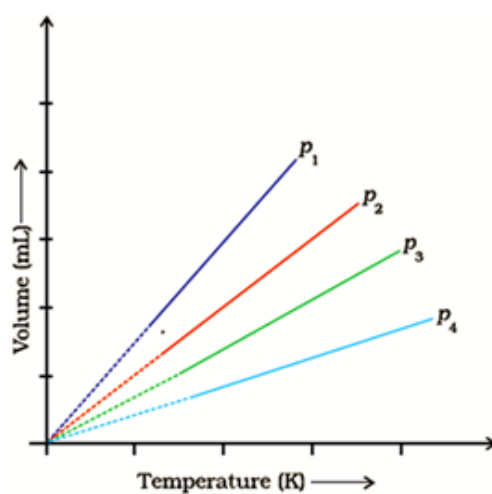


Fig. 5.1

- (i) $p_1 > p_2 > p_3 > p_4$
- (ii) $p_1 = p_2 = p_3 = p_4$
- (iii) $p_1 < p_2 < p_3 < p_4$
- (iv) $p_1 < p_2 = p_3 < p_4$

Answer: (iii)

Question: 4

The interaction energy of London force is inversely proportional to sixth power of the distance between two interacting particles but their magnitude depends upon

- (i) Charge of interacting particles
- (ii) Mass of interacting particles
- (iii) Polarisability of interacting particles
- (iv) Strength of permanent dipoles in the particles.

Answer: (iii)

Question: 5

Dipole-dipole forces act between the molecules possessing permanent dipole. Ends of dipoles possess 'partial charges'. The partial charge is

- (i) More than unit electronic charge
- (ii) Equal to unit electronic charge
- (iii) Less than unit electronic charge
- (iv) Double the unit electronic charge

Answer: (iii)

Question: 6

The pressure of a 1: 4 mixture of dihydrogen and dioxygen enclosed in a vessel is one atmosphere. What would be the partial pressure of dioxygen?

- (i) $0.8 \times 10^5 \text{ atm}$
- (ii) 0.008 Nm^{-2}
- (iii) $8 \times 10^4 \text{ Nm}^{-2}$
- (iv) 0.25 atm

Answer: (iii)