Chemistry Class - 11: Chapter – 1. Some Basic Concepts of Chemistry – Part-1

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I. Multiple Choice Questions (Type-I)

Questions 1:

Two students performed the same experiment separately and each one of them recorded two readings of mass which are given below. Correct reading of mass is 3.0 g. On the basis of given data, mark the correct option out of the following statements.

Student	Readings	
A	3.01	2.99
B	3.05	2.95
Student and Their Readings		

(i) Results of both the students are neither accurate nor precise.

(ii) Results of student A are both precise and accurate.

(iii) Results of student B are neither precise nor accurate.

(iv) Results of student B are both precise and accurate.

Answer: (ii)

Questions-2

A measured temperature on Fahrenheit scale is $200°F$. What will this reading be on Celsius scale?

(i) $40°C$

(ii) $94°C$

(iii) $93.3°C$

(iv) $30°C$

Answer: (iii)

Questions-3

What will be the molarity of a solution, which contains 5.85 g of $NaCl\,(s)\,per\,500\,mL$?

(i) $4\,mol\,L^{-1}$

(ii) $20 \, mol \, L^- 1$

(iii) $0.2 \, mol \, L^{-1}$

(iv) $2 \, mol \, L^{-1}$

Answer: (iii)

Questions-4

If 500mL of$a \, 5M$ solution is diluted to $1500 \, mL$, what will be the molarity of the solution obtained?

(i) $1.5M$

(ii) $1.66M$

(iii) $0.017M$

(iv) $1.59M$

Answer: (ii)

Questions-5

The number of atoms present in one mole of an element is equal to Avogadro number. Which of the following element contains the greatest number of atoms?

(i) $4g$He

(ii) $46g \, Na$

(iii) $0.40g \, Ca$

(iv) $12g$He

Answer: (iv)

Questions-6

If the concentration of glucose $(C_6 H_{12} O_6)$ in blood is $0.9 g L^{-1}$, what will be the molarity of glucose in blood?

(i) $5M$

(ii) $50M$

(iii) $0.005M$

(iv) $0.5M$

Answer: (iii)

Questions-7

What will be the molality of the solution containing $18.25 g$of HCl gas in$500g$ of water?

(i) $0.1m$

(ii) $1M$

(iii) $0.5m$

(iv) $_{1m}$

Answer: (iv)

Questions-8

One mole of any substance contains 6.022×10^{23} atoms/molecules. Number of molecules of H_2SO_4 present in $100mL$ of $0.02 M H_2SO_4$ solution is _____.

(i) 12.044×10^{20} molecules

(ii) 6.022×10^{23} molecules

(iii) 1×10^{23} molecules

(iv) 12.044×10^{23} molecules

Answer: (i)